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ANSWER Answer the

following questions in

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the space provided. 1.

_____ The coefficients
in a chemical

equation represent

the. (a) masses in

grams of all reactants
and products.

(b) relative number of
moles of reactants
and products.

(c) number of atoms of
each element in each
compound in a ...

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PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____

The actual yield of a

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reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2.

6.0 mol of N_2 are mixed with 12.0 mol of H

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SECTION 3

PROBLEMS Write the

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answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2$

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Chapter 3) deals with
the mass rela-

tionships of elements
in

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involves the mass

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reactants and products in a chemical reaction.

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conversion factor that
relates the amounts
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the substance that
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stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or

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